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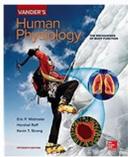
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CHAPTER 2 REVIEW QUESTIONS

- Describe the electric charge, mass, and location of the three major subatomic particles in an atom.  
*The electron has one unit of negative charge and revolves around the atom's nucleus in orbitals. The proton has one unit of positive charge and is found in the nucleus. The neutron has no charge and is also found in the nucleus.*
- Which four kinds of atoms are most abundant in the body?  
*Hydrogen (H), oxygen (O), carbon (C), and nitrogen (N).*
- Describe the distinguishing characteristics of the three classes of essential chemical elements found in the body.  
*The four major inorganic elements account for 99.9% of the total atoms in the body. There are seven essential mineral elements—calcium (Ca), chlorine (Cl), magnesium (Mg), phosphorus (P), potassium (K), sodium (Na), and sulfur (S). These are the most abundant elements in dissolved in the body fluids. Most of the body's inorganic and phosphate elements make up the solid matrix of bone tissue. There are other elements, called a total of 20 elements, are present in extremely small quantities but are nevertheless essential for several body functions, such as growth and the fluid transport of oxygen.*
- How many covalent bonds can be formed by atoms of carbon, nitrogen, oxygen, and hydrogen?  
*Carbon can form four, nitrogen three, oxygen two, and hydrogen one.*
- What property of covalence allows them to change their three-dimensional shape?  
*Molecules are not rigid structures. Atoms can rotate around their covalent bonds to form different shapes.*
- Define ionic and an ionic bond.  
*Ion: an atom that has either gained or lost one or more electrons from its outer electron shell, and thereby acquired a net charge.  
Ionic bond: a bond between electrically oppositely charged ions, such as Na<sup>+</sup> and Cl<sup>-</sup>.*

7. Draw the structure of an ionized carboxyl group and an ionized amino group.

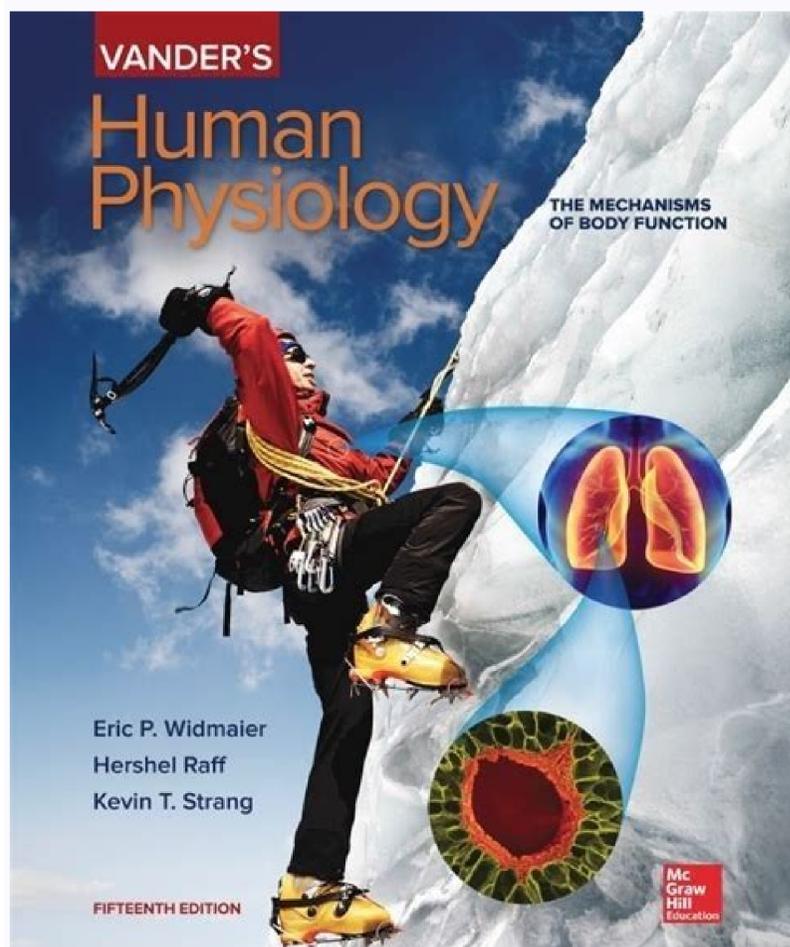


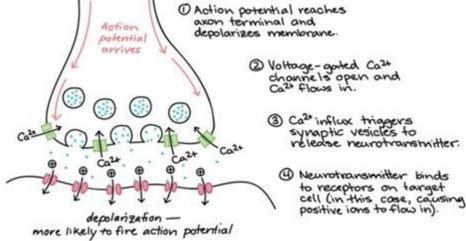
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